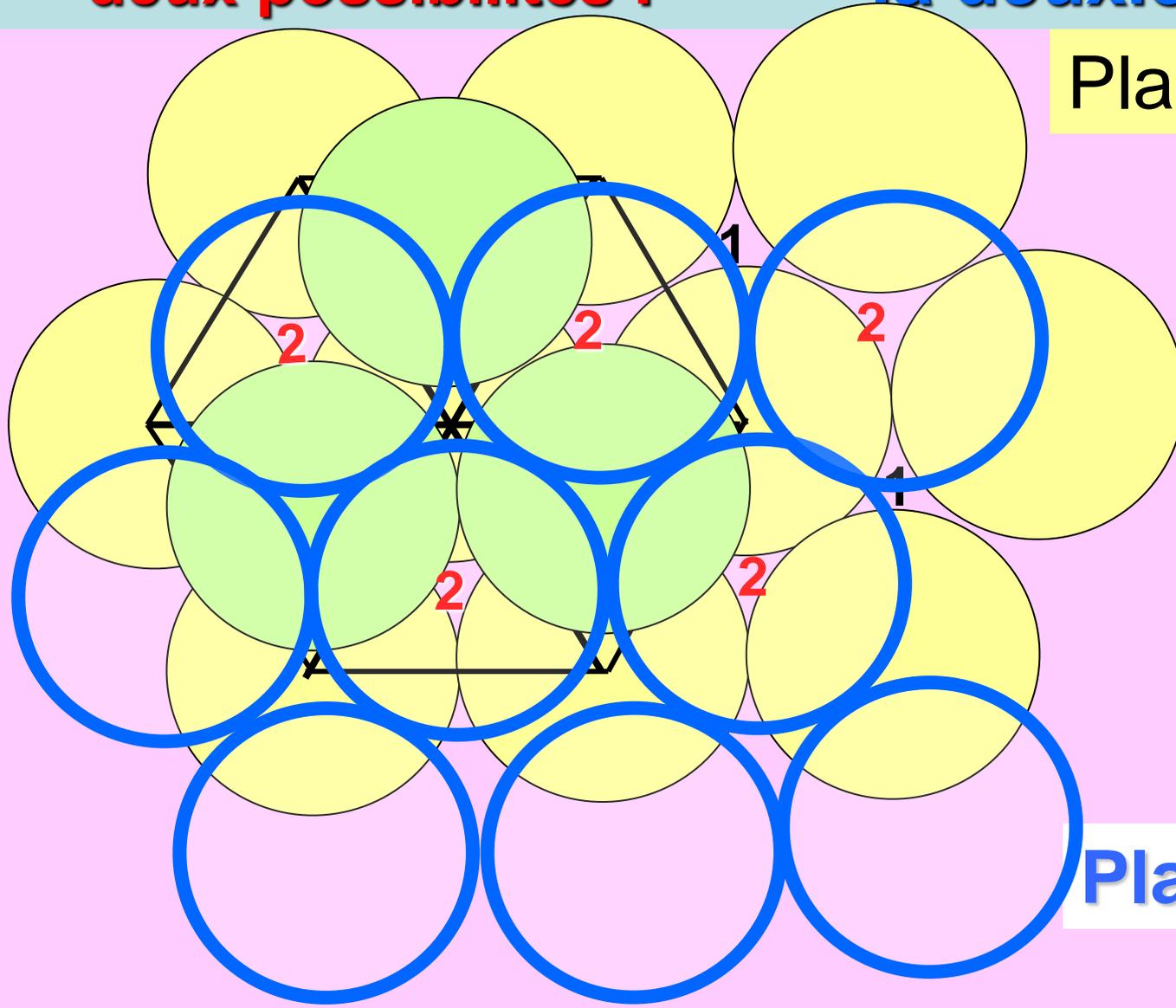


2- Troisième plan compact C :

deux possibilités :

la deuxième est

Plan A



Plan B

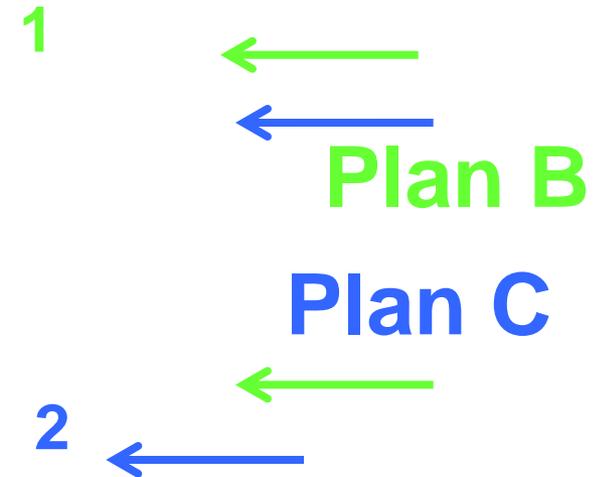
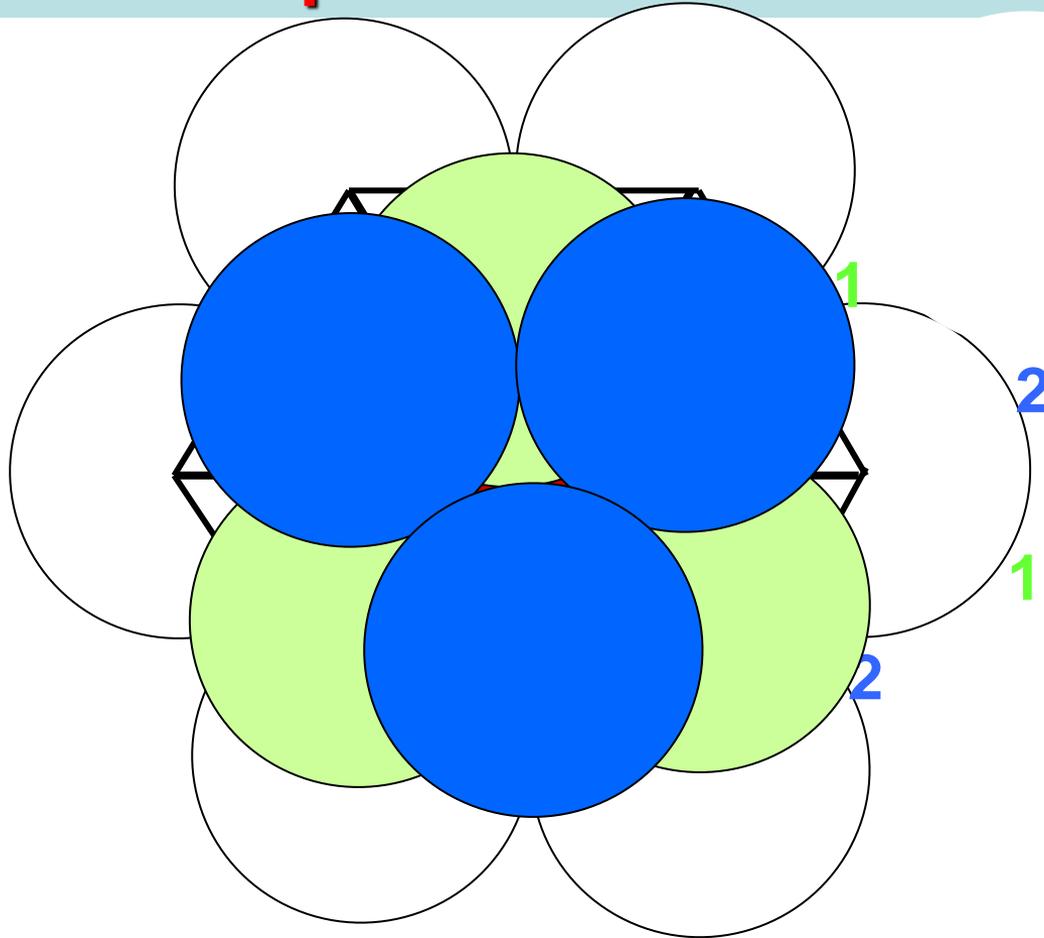
Plan C

2- Troisième plan compact

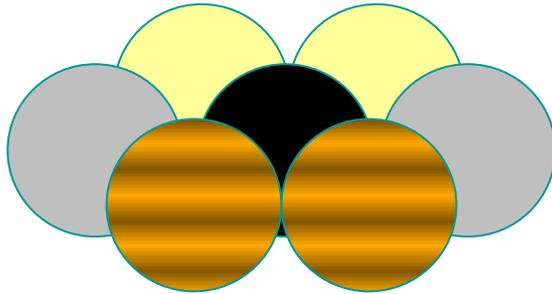
C :

deux possibilités

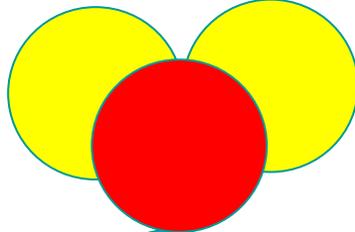
la deuxième est



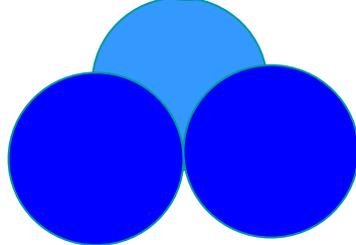
Plan D = A



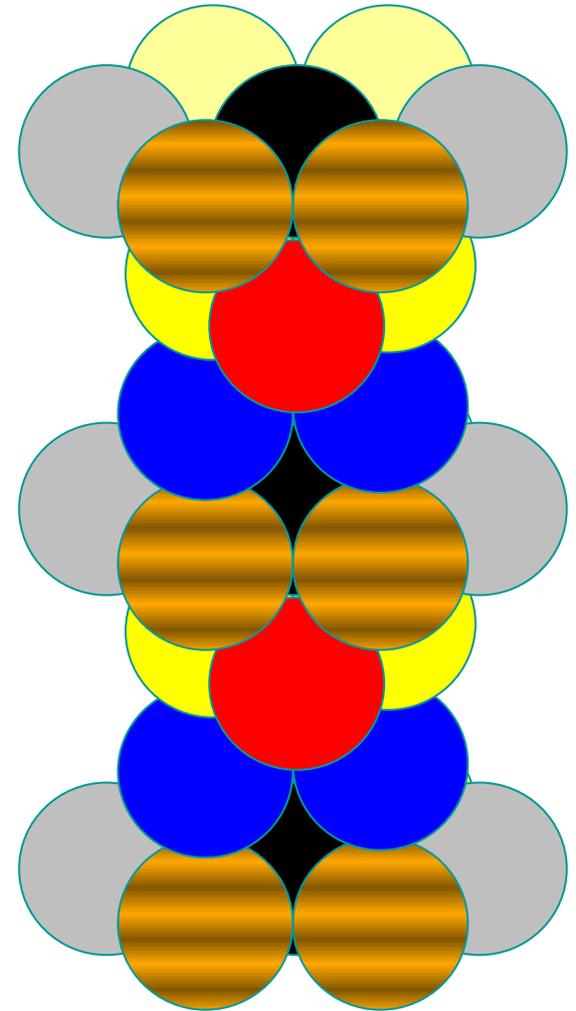
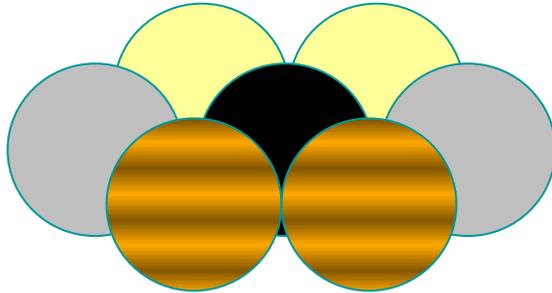
Plan C



Plan B



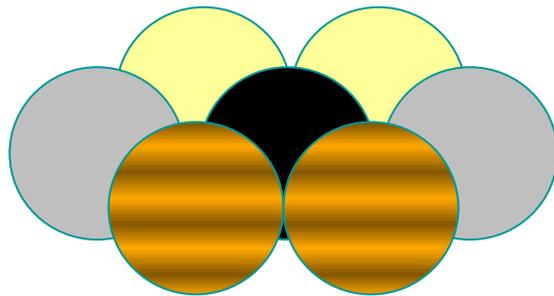
Plan A



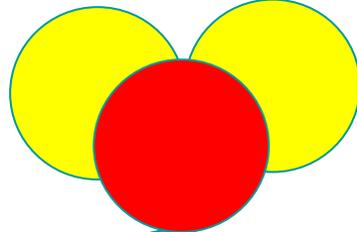
Succession des plans

A B C A B C A B

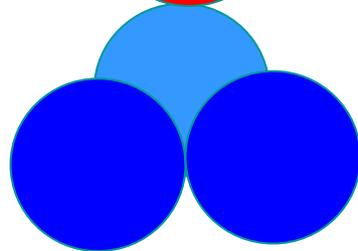
Plan D = A



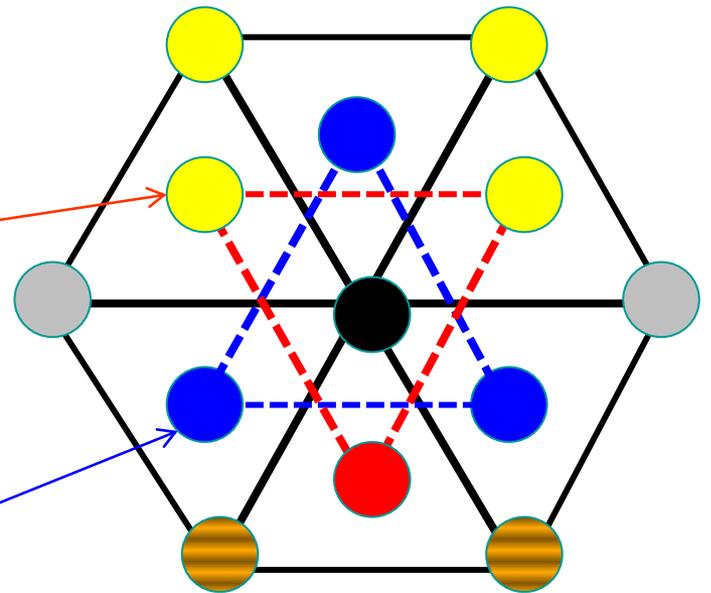
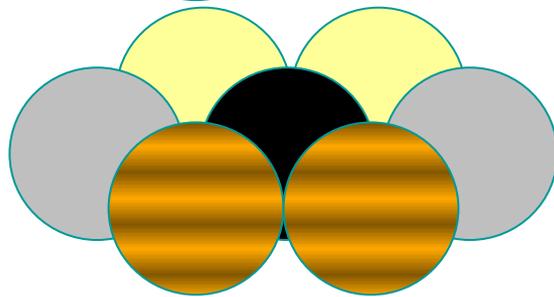
Plan C

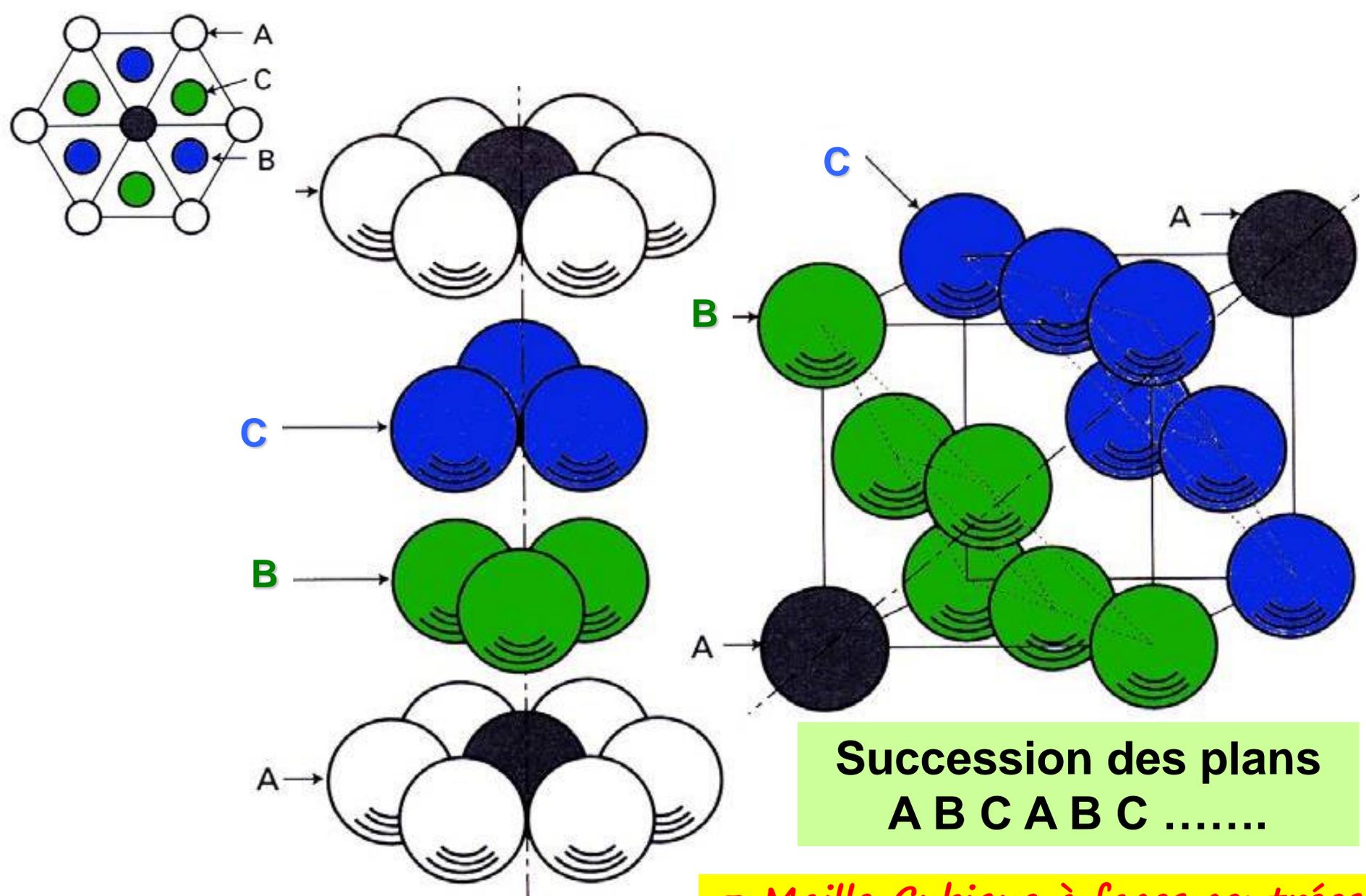


Plan B

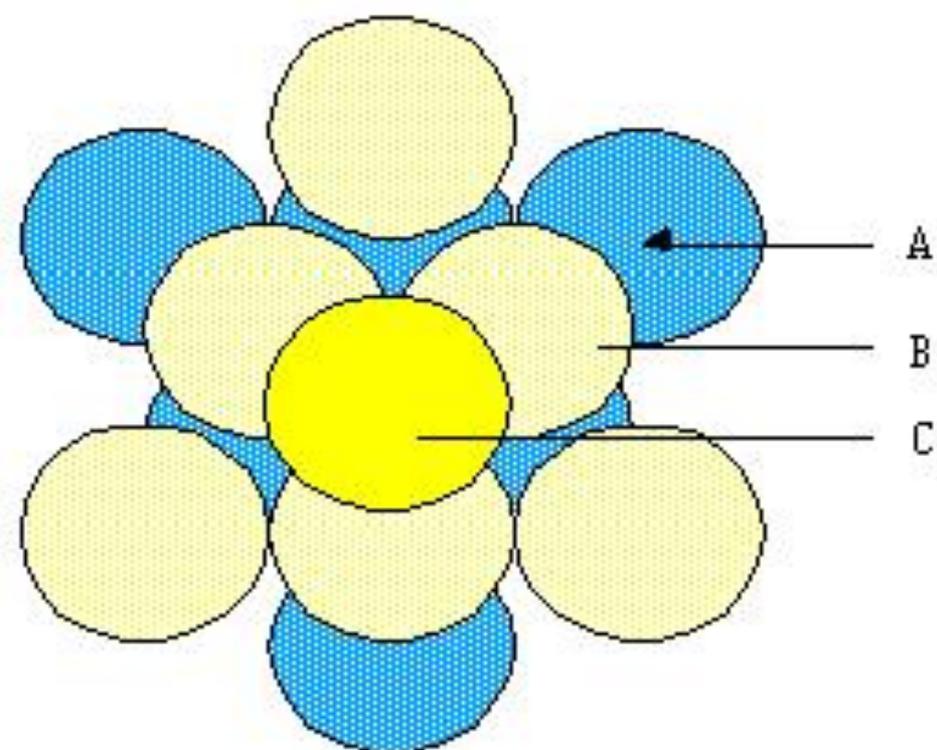


Plan A

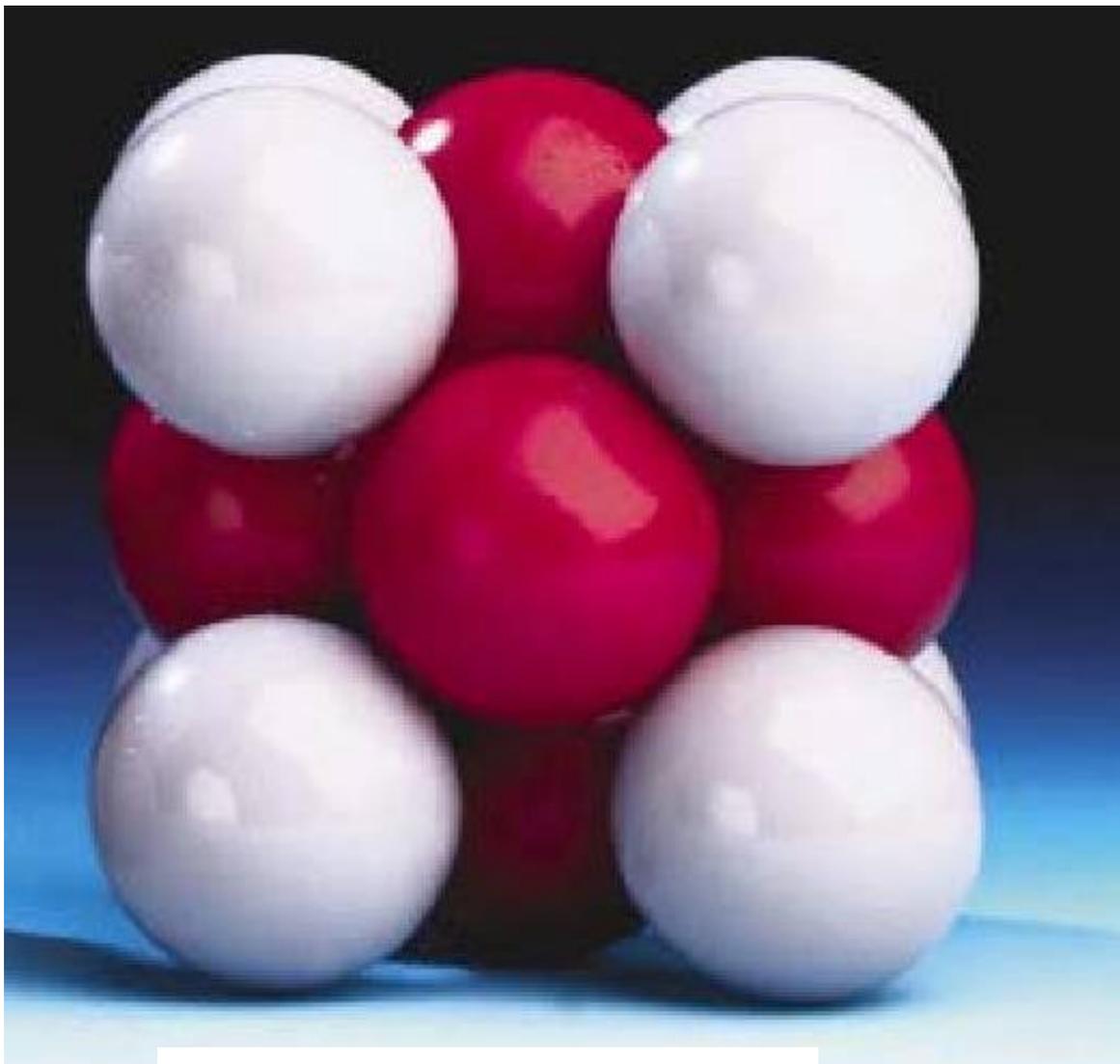




= Maille Cubique à faces centrées
CFC

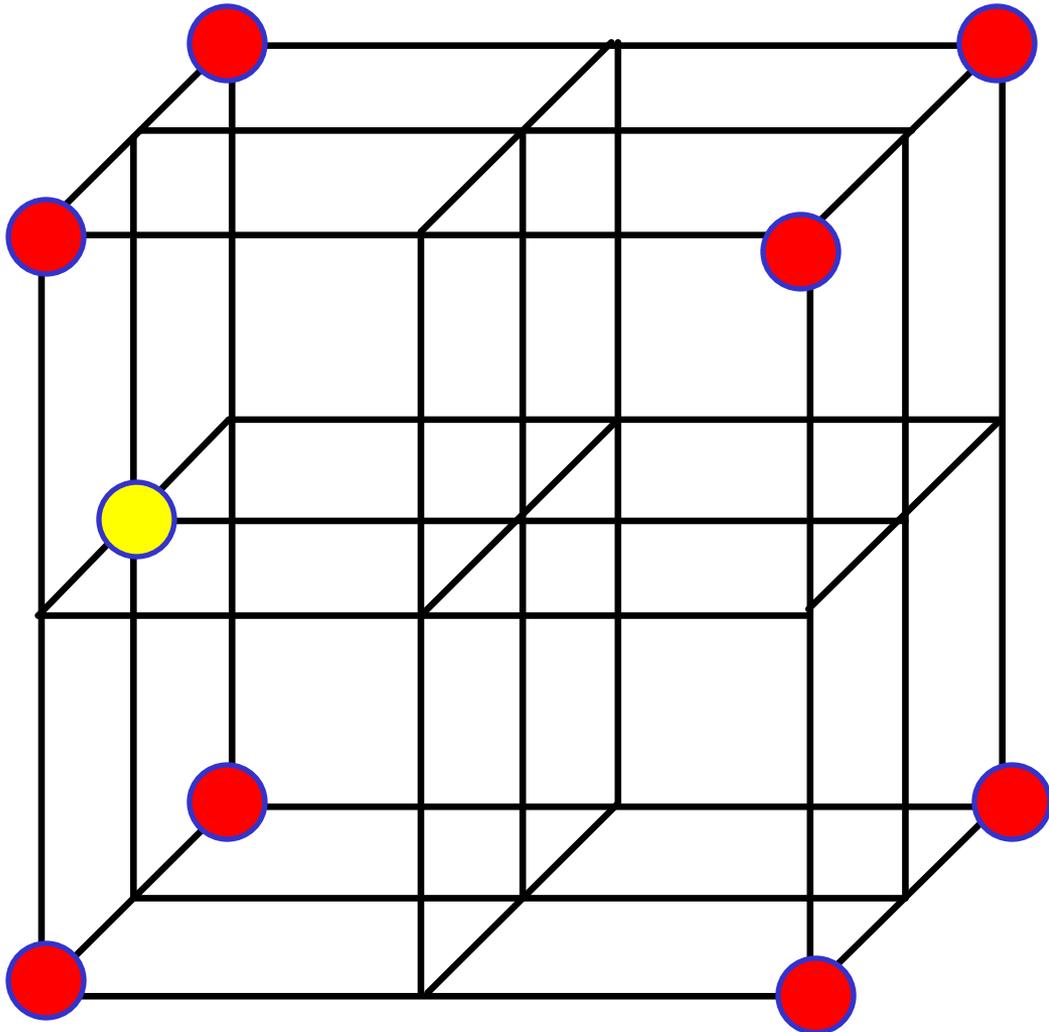


L'empilement compact ABCABC...
donne une structure cubique à faces centrées.



**Maille Cubique à
faces centrées**

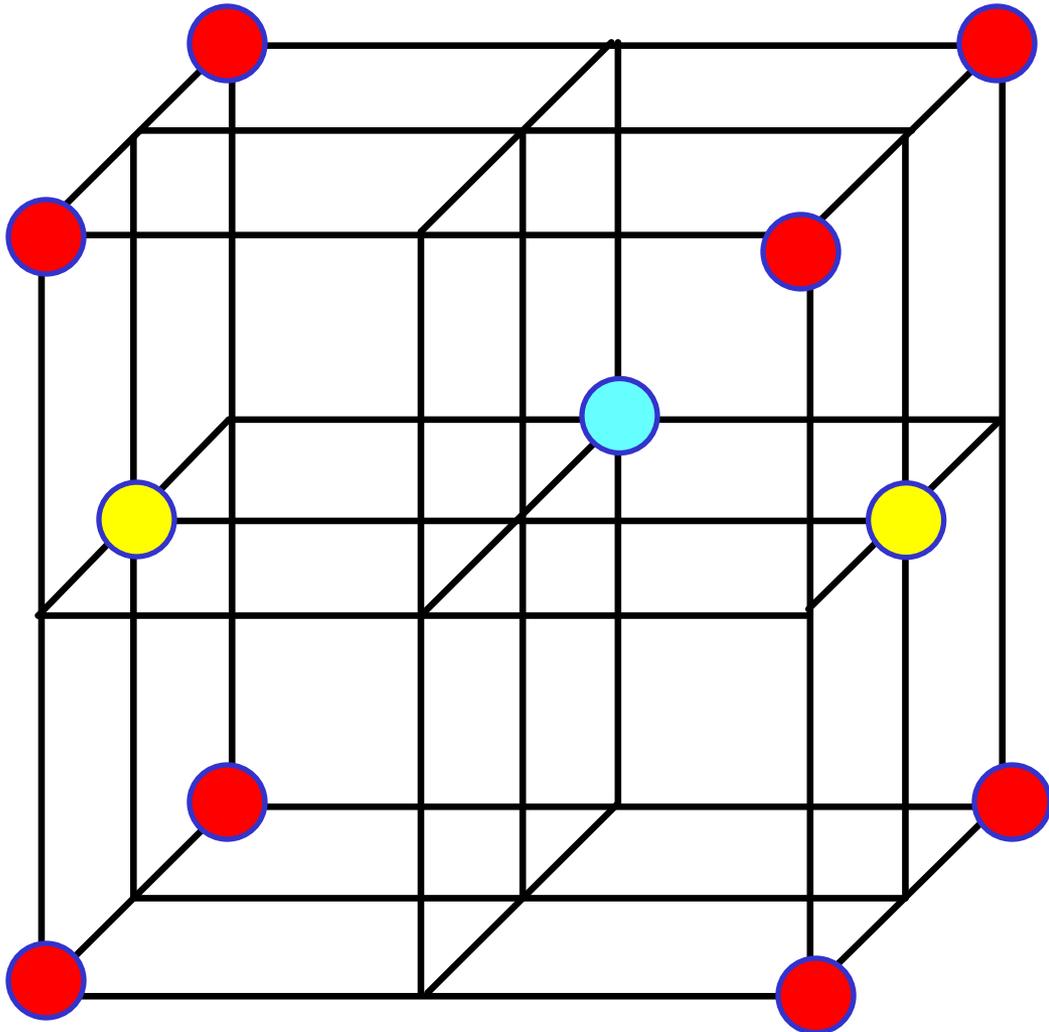
CFC



Coordonnées réduites

$(0, 0, 0)$

$(1/2, 0, 1/2)$

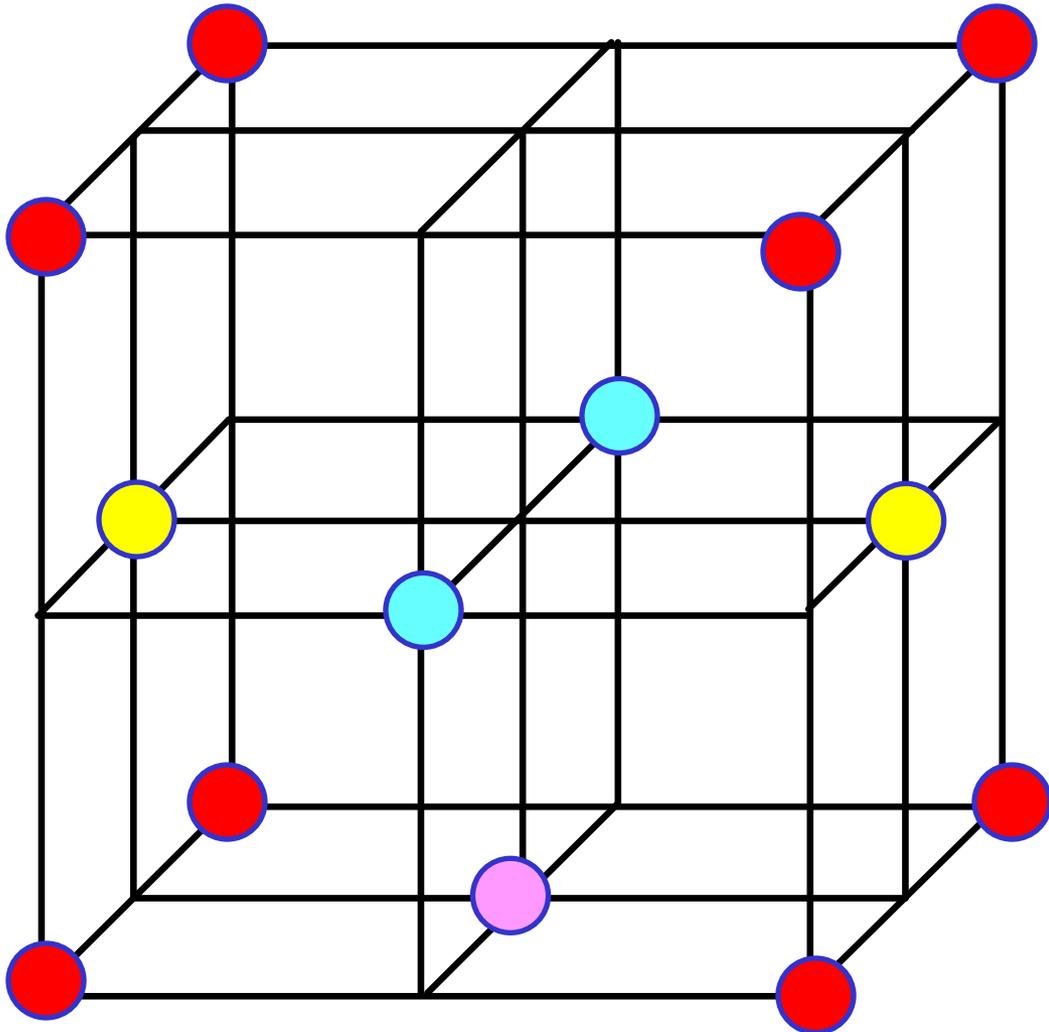


Coordonnées réduites

$(0, 0, 0)$

$(1/2, 0, 1/2)$

$(0, 1/2, 1/2)$



Coordonnées réduites

$(0, 0, 0)$

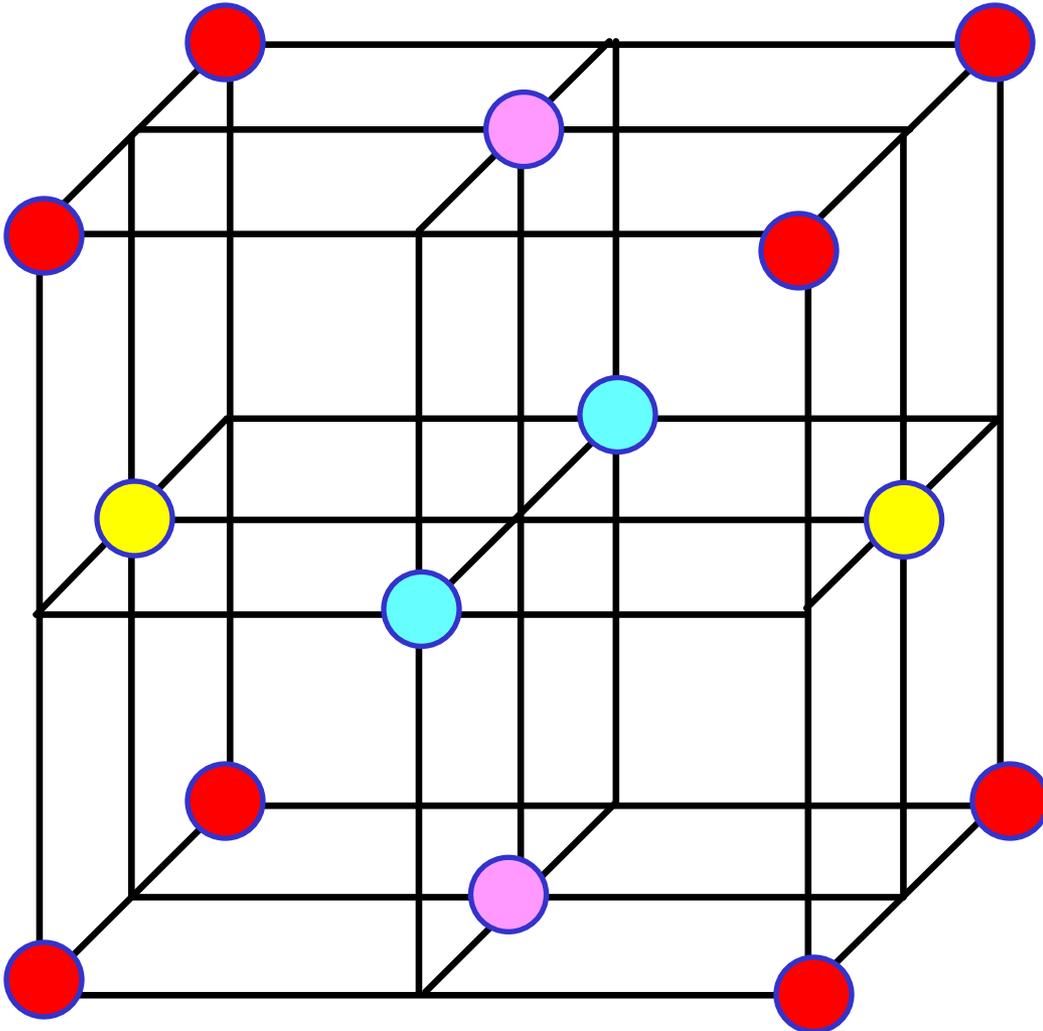
$(1/2, 0, 1/2)$

$(0, 1/2, 1/2)$

$(1/2, 1/2, 0)$

Structure Cubique à faces centrées

Mode du réseau cubique: Mode F



Coordonnées réduites

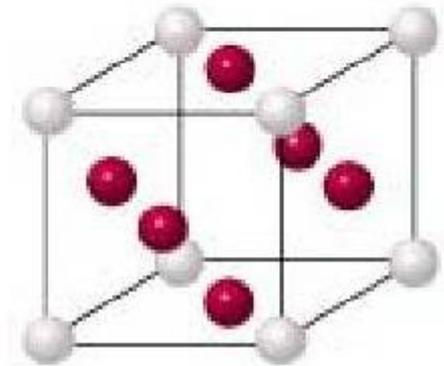
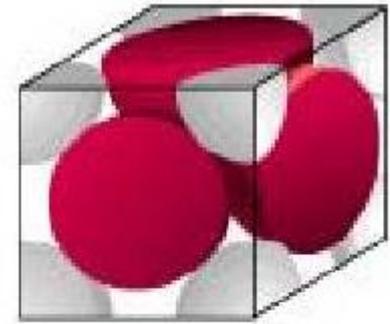
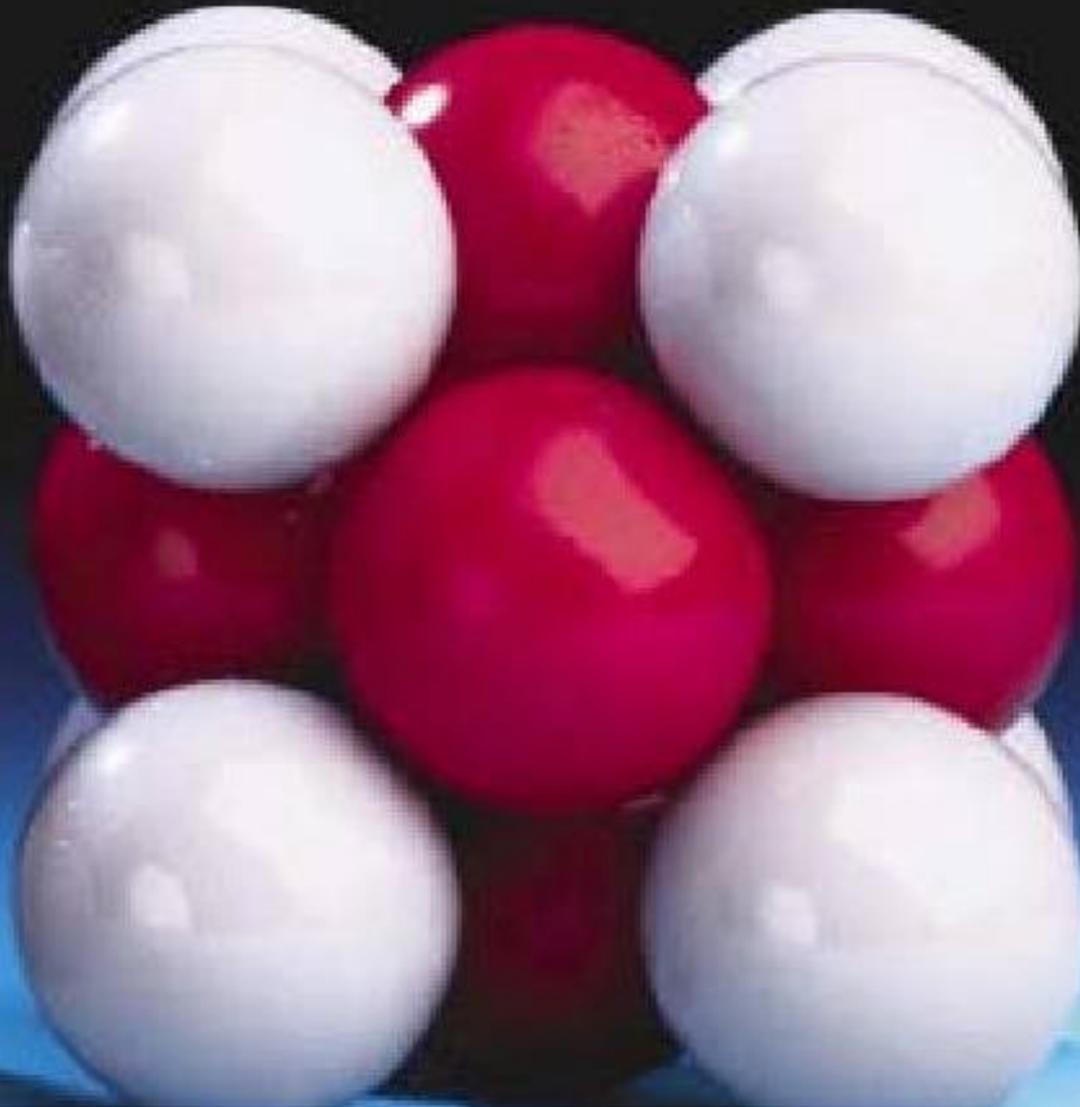
$(0, 0, 0)$

$(1/2, 0, 1/2)$

$(0, 1/2, 1/2)$

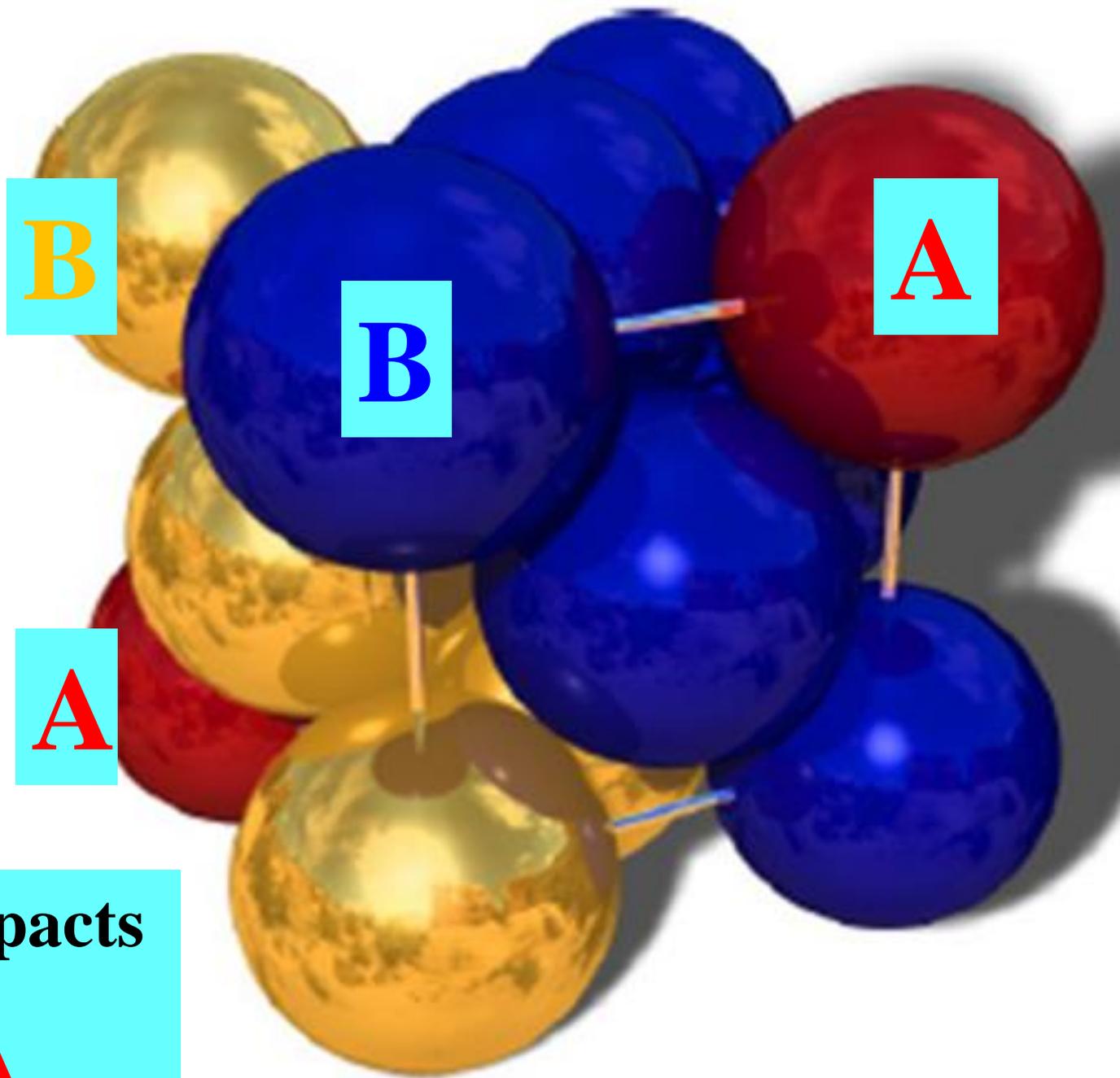
$(1/2, 1/2, 0)$

Structure Cubique à faces centrées, Mode F



On compte donc 8 atomes (sommets) \times $1/8$
+ 6 atomes (faces) \times $1/2$

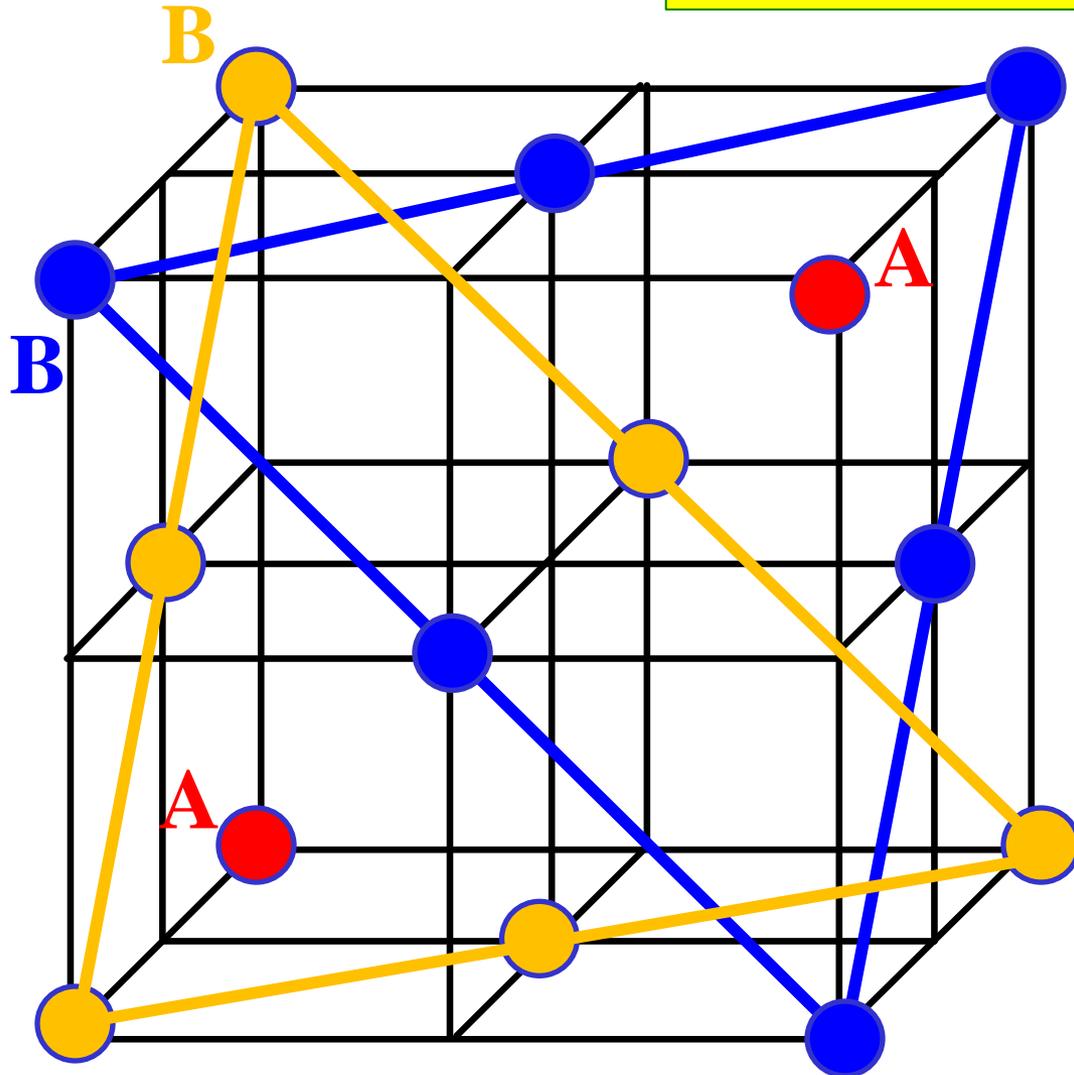
= 4 atomes/maille



Plans compacts
A B C A

Structure Cubique à faces centrées

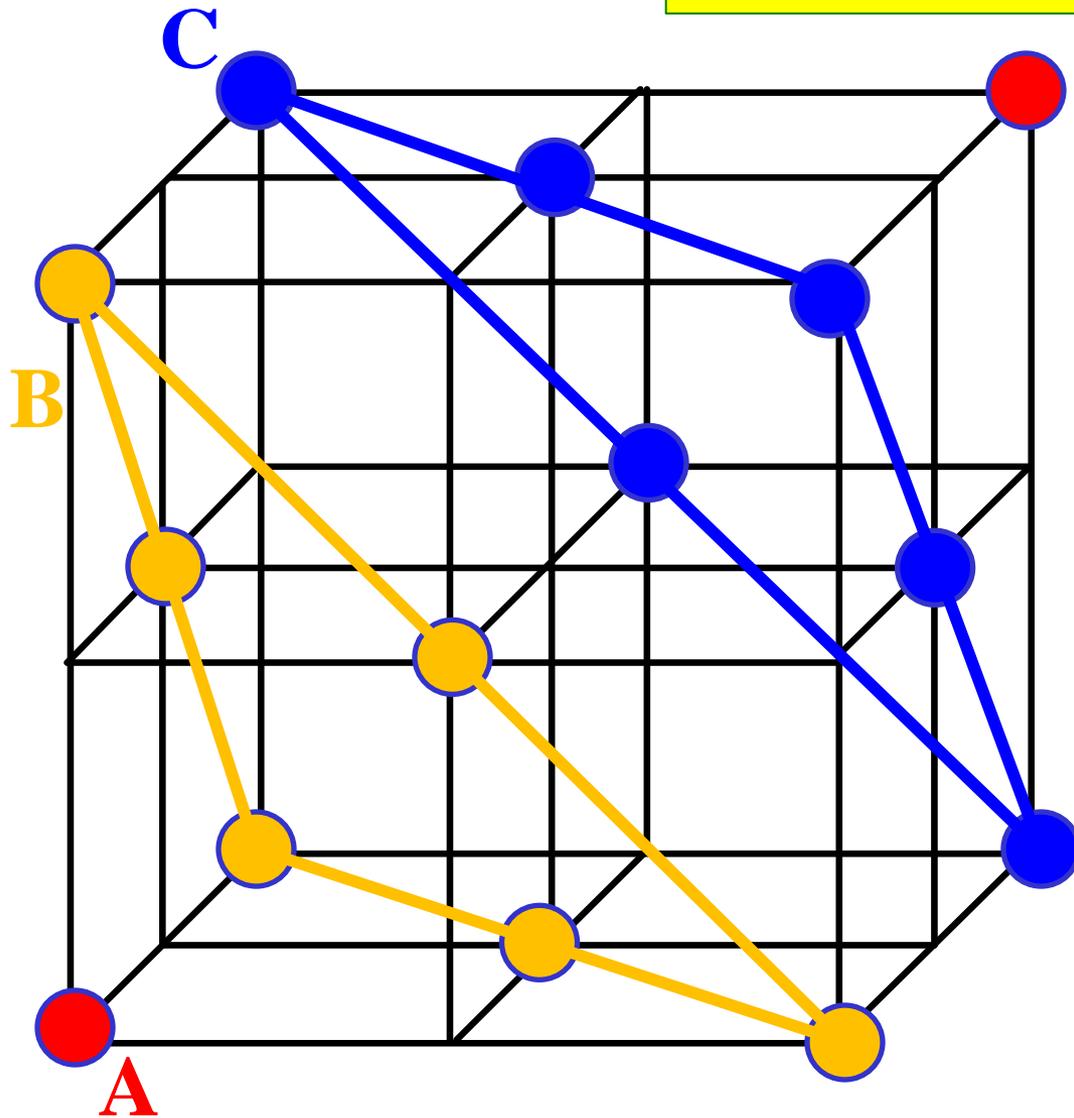
Mode du réseau cubique: Mode F



Plans compacts **ABCA**

Autre présentation : Structure Cubique à faces centrées

Mode du réseau cubique: Mode F



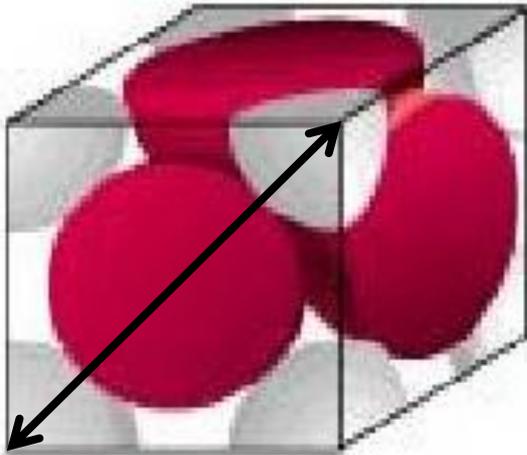
Plans compacts **ABCA**

Structure Cubique à faces centrées, Mode F

$\tau = n \cdot \text{volume}(1 \text{ atome}) / V(\text{maille})$ avec $n = 4$ atomes/maille

$$v(1 \text{ atome}) = \left(\frac{4}{3}\right) \pi R^3 \quad \text{et} \quad V(1 \text{ maille}) = a^3$$

Or pour un CFC, les atomes sont tangents selon la diagonale de la face

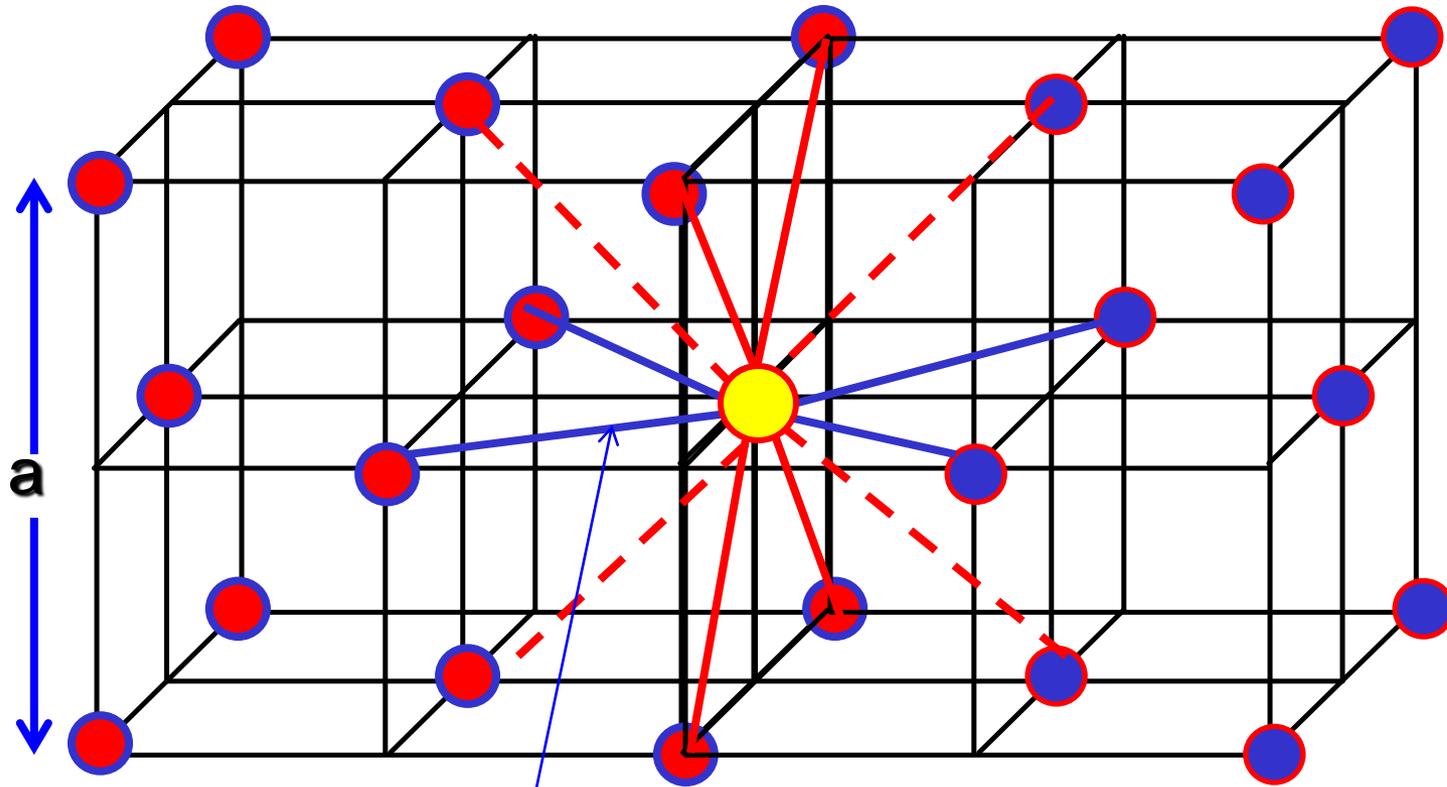


Relation de tangence :
 $4R = a\sqrt{2}$

D'où la relation :

$$\tau = \frac{4 \cdot \left(\frac{4}{3}\right) \pi R^3}{a^3} = \frac{4 \cdot \left(\frac{4}{3}\right) \pi \left(\frac{a\sqrt{2}}{4}\right)^3}{a^3} = 0,74$$

Structure Cubique à faces centrées, Mode F



coord. = ?

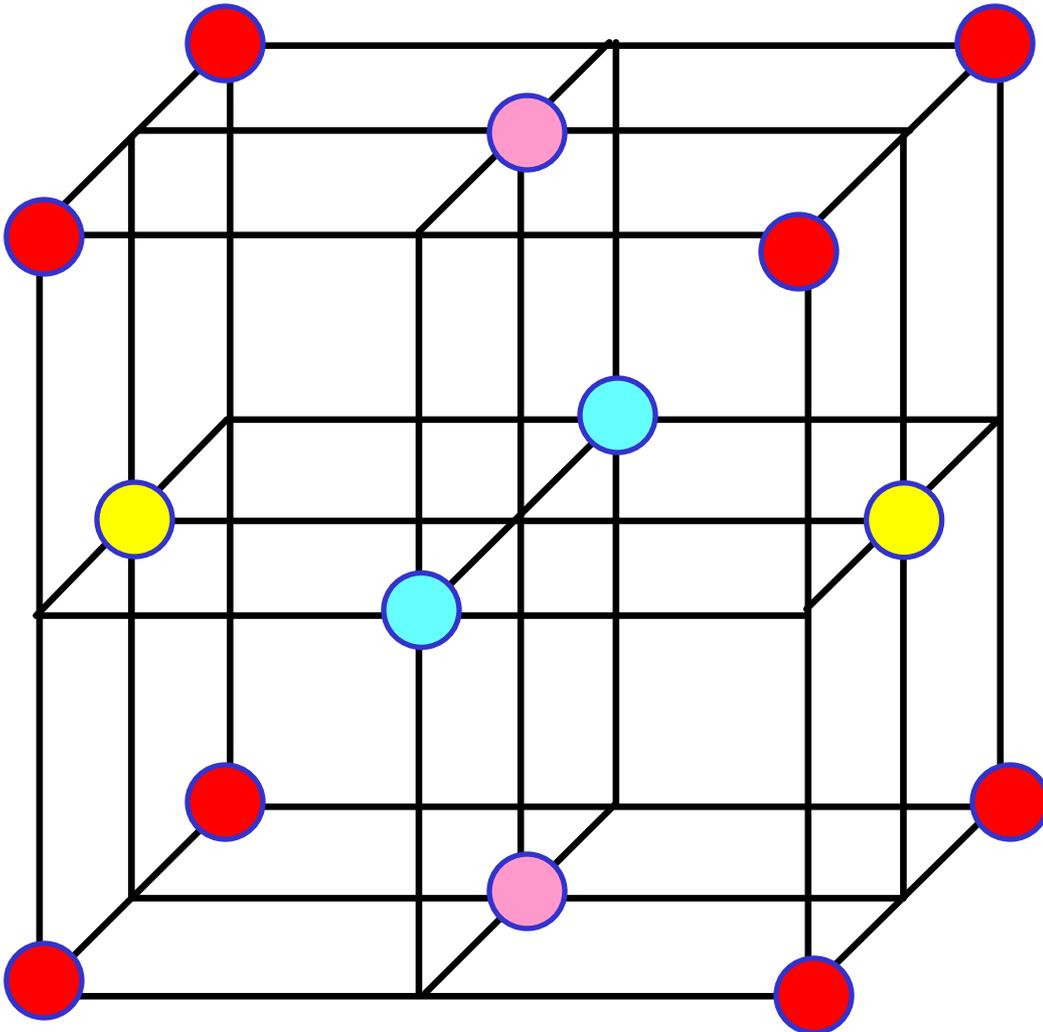
2 atomes sont situés à la distance $a\sqrt{2}/2$.

Chaque atome a 12 proches voisins

coord. = 12

Structure Cubique à faces centrées

Mode du réseau cubique: Mode F



Coordonnées réduites

$(0, 0, 0)$

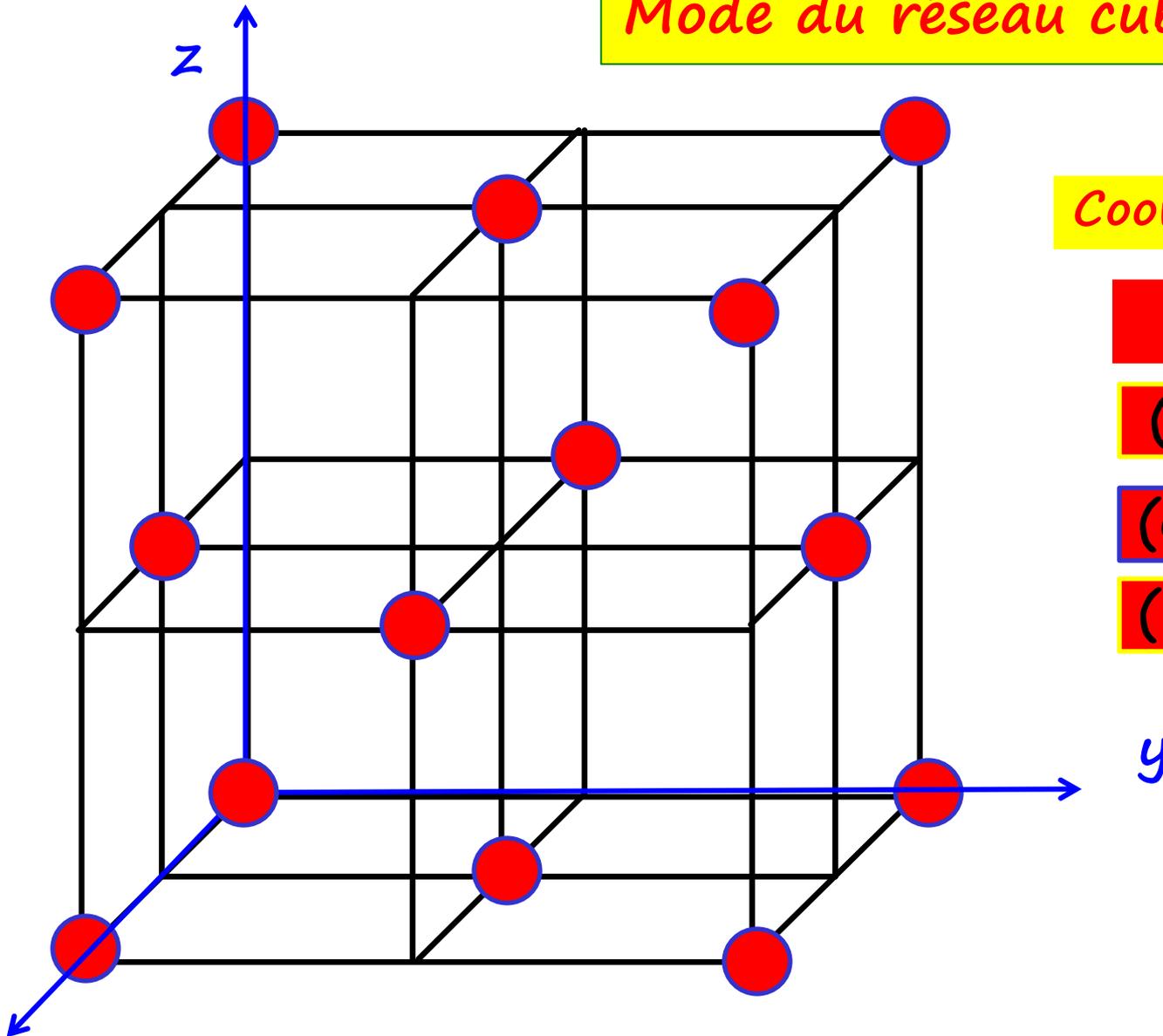
$(1/2, 0, 1/2)$

$(0, 1/2, 1/2)$

$(1/2, 1/2, 0)$

Structure Cubique à faces centrées

Mode du réseau cubique: Mode F



Coordonnées réduites

$(0, 0, 0)$

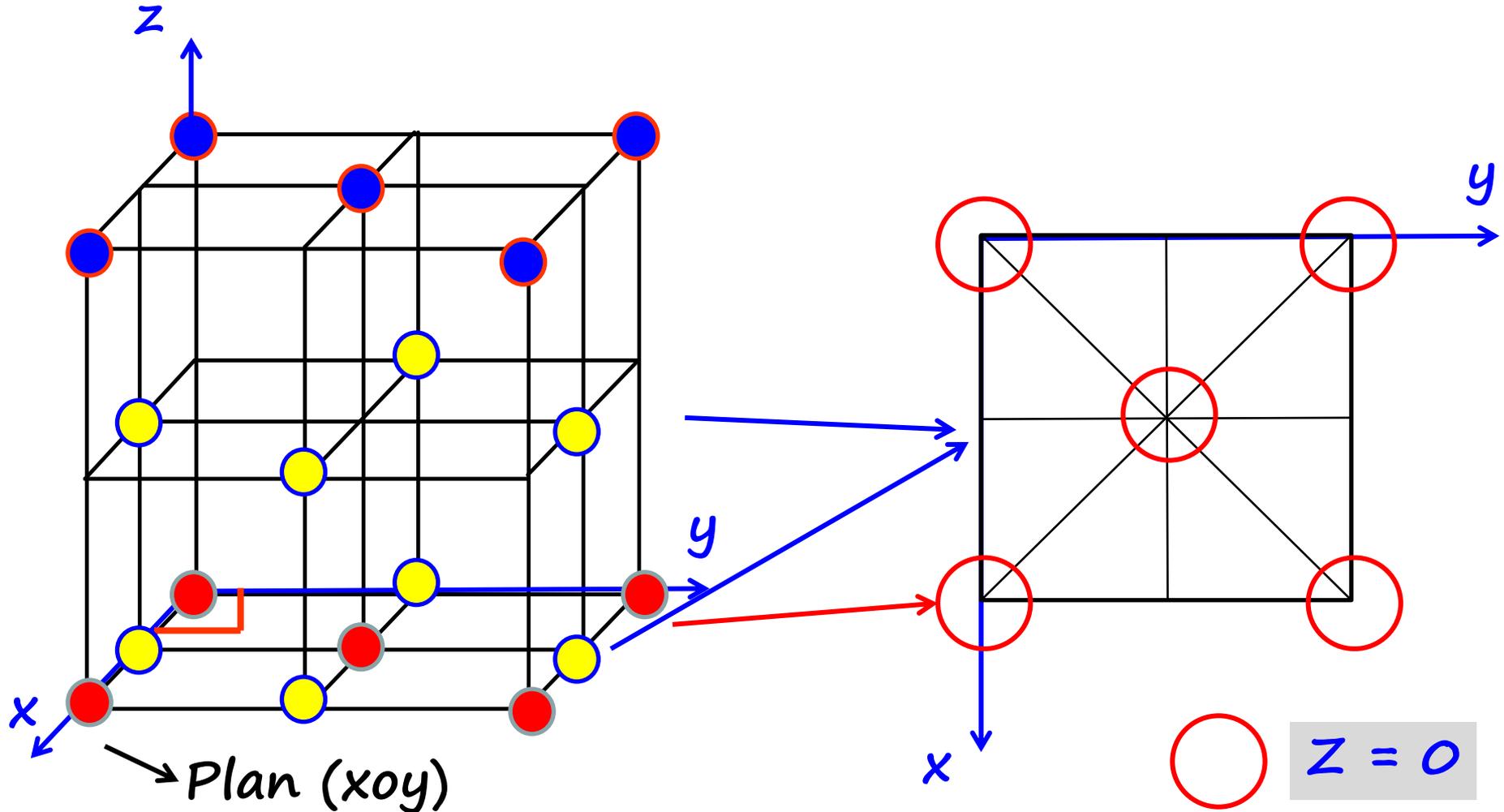
$(1/2, 0, 1/2)$

$(0, 1/2, 1/2)$

$(1/2, 1/2, 0)$

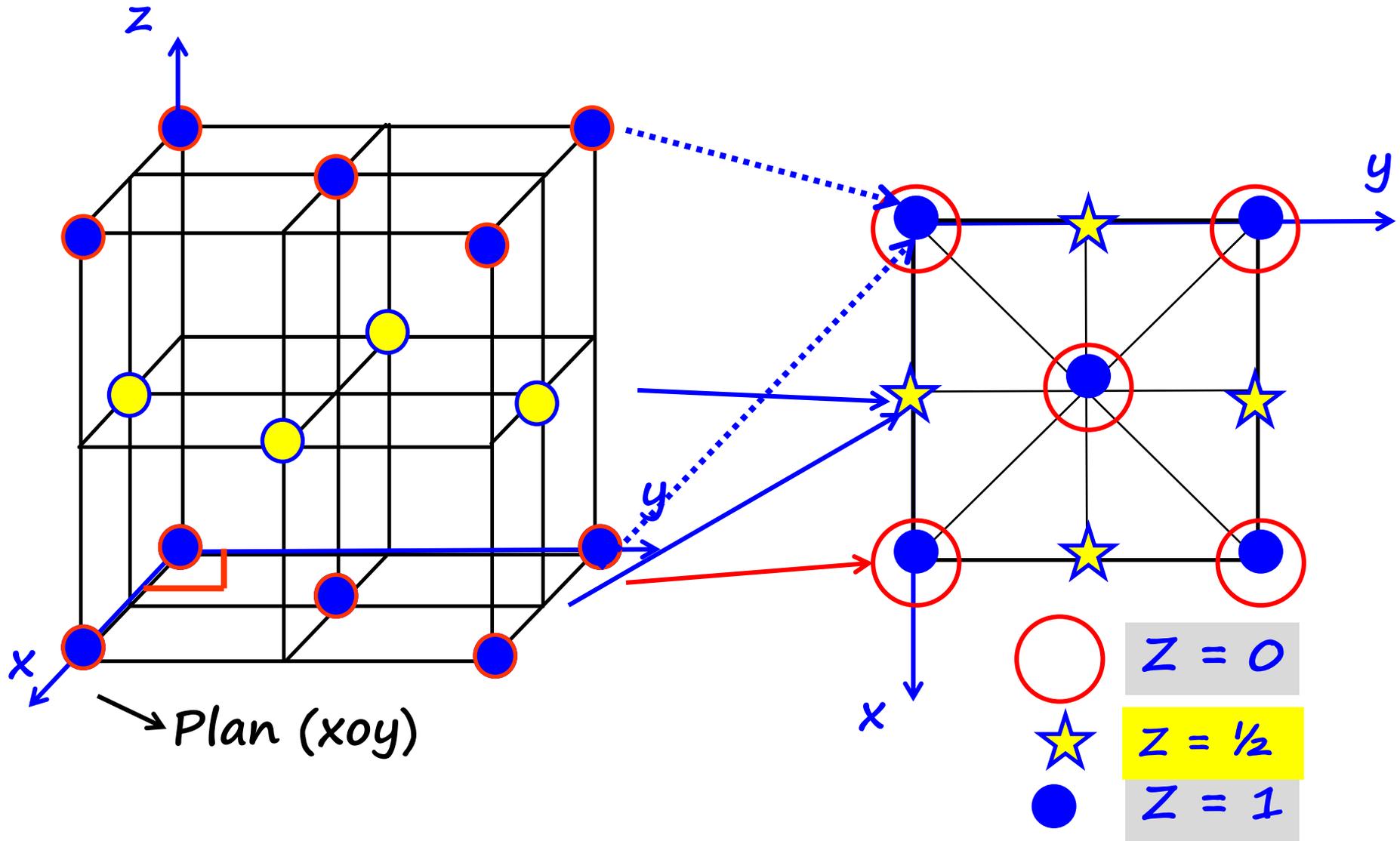
Structure Cubique à faces centrées, CFC

Mode du réseau cubique: Mode F



Structure Cubique à faces centrées, CFC

Mode du réseau cubique: Mode F

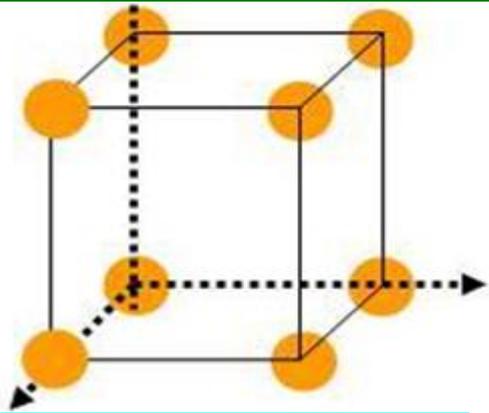


Résumé

Cubique simple, Mode P

Cubique centré, Mode I

CFC, Mode F



coord. = 6

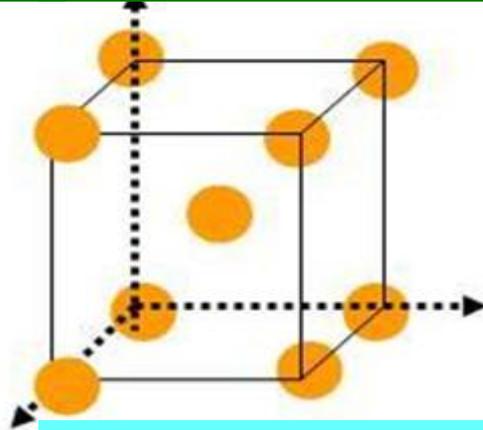
$\tau = 0,52$

1 at./maille

$2R = a$

Coordonnées réduites

$(0, 0, 0)$



coord. = 8

$\tau = 0,68$

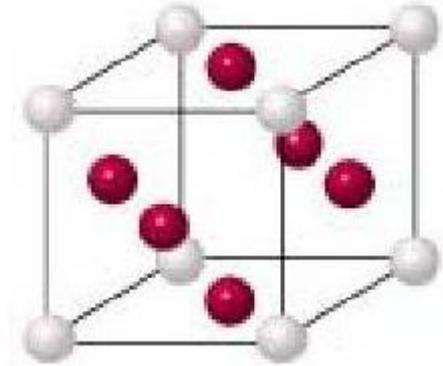
2 at./maille

$4R = a\sqrt{3}$

Coordonnées réduites

$(0, 0, 0)$

$(1/2, 1/2, 1/2)$



coord. = 12

$\tau = 0,74$

4 at./maille

$4R = a\sqrt{2}$

Coordonnées réduites

$(0, 0, 0)$

$(1/2, 1/2, 0)$

$(1/2, 0, 1/2)$

$(0, 1/2, 1/2)$